







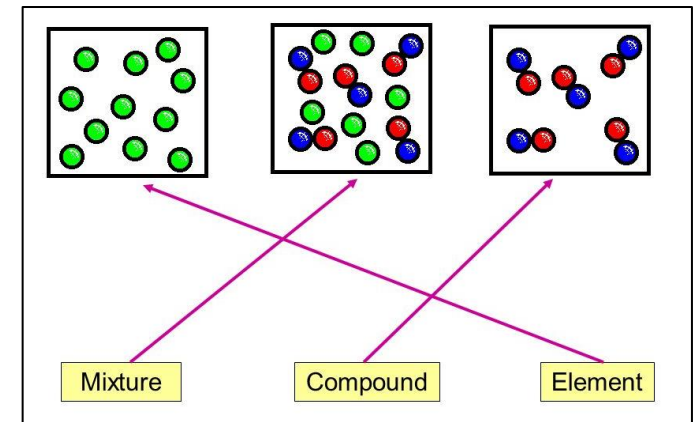
Key vocabulary	
compare and group	Putting objects together due to their similarities and differences
solids	Atoms closely packed, rigid structure
liquids	Atoms loosely packed, able to be poured
gases	Atoms move around freely
state	All matter exists as solids liquids or gases these are the different state
degrees Celsius (°C)	A temperature scale
dissolving	To become incorporated into a liquid
mixing	To combine or put together
reversible changes	Changes that can be undone or reversed
irreversible changes	Changes that cannot be undone or reversed
formation	An act of giving form or shape to something or of taking form
acid	A liquid that has lots of hydrogen ions
chemicals	A substance that is produced or used in a process (reaction) involving changes to atoms or molecules
atoms	The basic building block for all matter in the universe
elements	A pure substance that is made from a single type of atom.
compounds	Two or more different elements chemically bonded together.
Periodic table	A system for arranging the chemical elements
reactions	A process in substances are converted to different substances.
dilute	To make thinner by adding more liquid
universal indicator	A test to measure the pH value of a liquid

Chemical Reactions

Year 6

Reversible Changes	Irreversible Changes
 <p>Melting ice cubes</p>	 <p>Burning wood</p>
 <p>Freezing water</p>	 <p>Making bread</p>











Significant scientists	
<p>Marie Curie (1867-1943)</p> 	<p>Marie Curie is most famous for the discovery of the elements polonium and radium. Prohibited from higher education in her native Poland (then controlled by Russia), she moved to Paris in 1891 and studied at the Sorbonne.</p>
<p>Dmitri Mendeleev (1834-1907)</p> 	<p>Dmitri was a Russian chemist who developed the periodic classification of the elements. Mendeleev found that, when all the known chemical elements were arranged in order of increasing atomic weight, the resulting table displayed a recurring pattern, or periodicity, of properties within groups of elements.</p>



Periodic Table of the Elements

1 H	H hydrogen																2 He																		
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne																		
11 Na	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar											19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe																		
55 Cs	56 Ba	57 La	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn																		
87 Fr	88 Ra	89 Ac	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og																		
lanthanoid series		58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tm	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu																				
actinoid series		90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr																				

Hazard Symbols

 Explosive H19A-S	 Oxidising H22A-S	 Extremely flammable H20A-S	 Corrosive H18A-S	 Dangerous for the environment H21A-S
 Harmful H15A-S	 Highly flammable H13A-S	 Toxic H16A-S	 Irritant H14A-S	 Very toxic H17A-S